

## Chemquest 5 Specific Heat Answer Key

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### Chemquest 5 Specific Heat Answer

13. A beaker containing 250.0 g of water is heated with 1500.0 J of heat energy. If the temperature of the water changed from 22.0000oC to 23.4354oC, what is the specific heat of water?  $\Delta T = 23.4354 \text{ oC} - 22.0000 \text{ oC} = 1.4354 \text{ oC}$  14. Heat energy equal to 25,000 J is applied to a 1200 g brick whose specific heat is 2.45 J/goC.

### Critical Thinking Questions

: Specific Heat In questions 1 and 2 above you probably recognized that the temperature change of a substance depends on the mass of the substance. You also have probably experienced the fact that different substances heat at different rates as was discussed in questions 3 and 4 above.

### ChemQuest 5

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Critical Thinking Questions 8. Which substance has a higher specific heat—water or a metal like the aluminum in a car hood Since it requires more heat energy change the temperature of water and since specific heat is a measure of how much energy is needed to change the temperature of a substance, water must have a higher specific heat. 9.

### 16 ChemQuest 5 Name: - Yumpu.com

This is video made for a chem project and it is intended to educate other on specific heat.

### Chemquest 5:Specific Heat

16 ChemQuest 5 Name: \_\_\_\_\_ Date: \_\_\_\_\_ Hour: \_\_\_\_\_ Information: Heat and Temperature When a substance absorbs heat energy, the temperature of the substance increases. There are a number of factors (such as mass of the substance) that affect how much the temperature of a substance changes. Critical Thinking Questions 1. Consider two pots of water.

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UNIT 5 - HW Practice Keys - Mini-Lab Specific Heat Question #1 - ChemQuest 5: Specific Heat (First Pages) - ChemQuest 5: Specific Heat Skills (Last Page) - ChemQuest 53: 1st Law (First Pages) - ChemQuest 53: 1st Law Skills (Last Page)

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5Q?5 = all 4. Calculate the heat capacity of a piece of wood if 1500.0 g of the wood absorbs 6 104 joules of heat, and its temperature changes from 32oC to 57oC. 0 14 x (move (A 500 = (1500) 'c 61-32) 5. 100.0 of 4.00C water is heated until its temperature is 37 cc. If the specific heat of water is 4.18 J/g0C, calculate the amount of heat energy needed to cause

### Specific Heat Wksht20130116145212867

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Chemquest 10 - Specific Heat - 41 ChemQuest 10 Name Date Hour Information Heat and Temperature When a substance absorbs heat energy the temperature of. ...

Specific\_Heat\_Worksheet\_Answers Hazel Park High School SCIENCE 101 - Fall 2015

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## Chemquest 9 Average Atomic Mass Answer Key

Example answer: The specific heat of the hot chocolate was 3.9 J/g °C, which is less than the specific heat of water. One possible reason for the difference is the specific heat of the hot chocolate powder that was mixed into the water. The powder might have lowered the specific heat. Another possible reason for the difference could be ...

## Specific Heat Capacity Handout Answer Key

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Q. Specific heat of water is 4.18 J g<sup>-1</sup> ° C<sup>-1</sup>. Specific heat of wood is 1.760 J g<sup>-1</sup> ° C<sup>-1</sup> What material needs more heat energy to raise the temperature? answer choices

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5 significant figures Final rounded answer should have 3 significant figures since 3 is the least number of significant figures in this problem. 325.714285714 330 14 4560 = = 3 significant figures 2 significant figures Final rounded answer should have only 2 significant figures since 2 is the least number of significant figures in this problem.

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